Chemistry assignments for 5/4 – 5/8

Hey! : )

I hope you are all settling in with the work being given. If you haven’t submitted the work from last week please turn it in.

I just want to say that I am very proud of the amount of students participating in the doing the work even though you have credit! Great Job!!!! This will certainly benefit you down the road … if nothing else maybe it will keep you from getting Alzheimer’s since you are using your brain!

I know I said I would post your work on TEAMS last week but I didn’t get to that because of the amount of student/parent communication I was working on. Sorry about that. The weebly page seems to be working so I will stick to that unless people feel teams is that much better so give me some feedback concerning those two platforms. Which do you think is better, the way I have it or going to TEAMS? Thanks : )

**PLEASE READ**: You can submit your work before Friday but PLEASE submit your work all at once. It is much easier for me to grade it all together and place it in PowerSchool. Thanks : ) It also helps to label your work at the top of the paper like “ balancing equation ws 1”. When I send out the mass email on Monday please don’t “Reply to All” this happened several times last week. Make sure you select “Reply” which will only go to me or create a new email to me: [jschleder@richmond.k12.mi.us](mailto:jschleder@richmond.k12.mi.us)

This is new material but this week is pretty easy. Next week we will apply what you learn this week to a difficult topic of predicting products in chemical reactions. I won’t schedule any Microsoft Teams meetings this week but I will next week as we begin working on predicting products.

\* If you would like to review the work you have completed I put the answer keys for Balancing equation worksheet 1 and from Quiz 8-1. These are on my weebly page. Next week I will post the answers for 8-2 rev ws 1, 2, and 3 .

Monday 5/4: Vieiw the Slides for section 8-2 from the Chapter 8 PowerPoint and Read through the 8-2 Notes. Both files are attached to my web page <https://schleder.weebly.com> . The notes I make are a summary of what is in the text book.

This week I want you to concentrate on learning the 5 types of chemical reactions. Next week we will tackle a very difficult topic using the 5 types of reactions which is predicting the products of a chemical reaction.

Tuesday 5/5: Watch the video: Types of Chemical Reactions <https://www.youtube.com/watch?v=aMU1RaRulSo>

Write a 4 sentence summary about the video.

This is by Tyler DeWitt. There are a lot of videos out there on this topic. I think this guy does the best job (he has about 1.5 million views) but feel free to check out any of the other videos on YouTube. I think he really helps clarify the 5 types of reactions.

Wednesday 5/6: Do Chap 8-2 rev ws 1 (attached to weebly).

Please read the info below … these are things that have helped students in past years.

This covers Synthesis reactions. I’m not sure if the video above said it

or not but a synthesis reaction ALWAYS has **ONE** Product (products are the things to the right of the

arrow). Remember you don’t need to print the worksheets, just write your answers on a piece of paper.

Thursday 5/7: Do Chap 8-2 rev ws 2 (attached to weebly).

This covers Decomposition reactions. I’m not sure if the video above said it

or not but a decomposition reaction ALWAYS has **ONE** Reactant (reactants are the things to the left of the arrow). Also, you need to think about Synthesis and Decomposition reactions as being opposite of each

other.

Friday 5/8: Do Chap 8-2 rev ws 3 (weebly).

This covers Single-Displacement Reactions. I also use the word Replacement which

means the same thing as displacement. In the notes it says they often happen in an AQUEOUS solution

so that is one big clue to identifying them. Synthesis, Decomposition, and Combustion reactions never

happen in aqueous solutions. The video does a good job of using the “dance” analogy to help you

understand how single-displacement reactions work. I like to use boyfriend or girlfriend … dumping

one for a new one ; ) . One thing I don’t think the video points out is the single element can also be a

negative anion so therefore it will replace the negative anion. Bottom line is if the single element is a

positive cation it replaces the positive cation in in the compound. If it is a negative anion it will replace

the negative anion in the compound. Remember, in a chemical formula the positive cation ALWAYS is

written first and the negative anion is ALWAYS written second! Ex: Ca+2  OH -1  = Ca(OH)2

Please submit: 4 sentence video summary, 8-2 rev ws 1 answers, 8-2 rev ws 2 answers, 8-2 rev ws 3 answers