

COMMON IONS AND OXIDATION NUMBERS

1+	2+	3+	4+
ammonium NH_4^+	barium Ba^{2+}	aluminum Al^{3+}	lead(IV) Pb^{4+}
copper(I) Cu^+	cadmium Cd^{2+}	bismuth Bi^{3+}	tin(IV) Sn^{4+}
hydrogen H^+	calcium Ca^{2+}	chromium(III) Cr^{3+}	
lithium Li^+	copper(II) Cu^{2+}	iron(III) Fe^{3+}	
mercury(I) Hg_2^{2+}	iron(II) Fe^{2+}	nickel(III) Ni^{3+}	
potassium K^+	lead(II) Pb^{2+}	Cobalt(III) Co^{3+}	
silver Ag^+	magnesium Mg^{2+}		
sodium Na^+	mercury(II) Hg^{2+}		
	nickel(II) Ni^{2+}		
	strontium Sr^{2+}		
	tin(II) Sn^{2+}		
	zinc Zn^{2+}		
	Cobalt(II) Co^{2+}		

1-	2-	3-
acetate $\text{C}_2\text{H}_3\text{O}_2^-$	carbonate CO_3^{2-}	phosphate PO_4^{3-}
bromide Br^-	chromate CrO_4^{2-}	phosphide P^{3-}
chlorate ClO_3^-	dichromate $\text{Cr}_2\text{O}_7^{2-}$	nitride N^{3-}
chlorite ClO_2^-	oxalate $\text{C}_2\text{O}_4^{2-}$	OTHERS
chloride Cl^-	oxide O^{2-}	bicarbonate HCO_3^-
cyanide CN^-	peroxide O_2^{2-}	bisulfate HSO_4^-
fluoride F^-	sulfate SO_4^{2-}	hydrogen phosphate HPO_4^{2-}
hydride H^-	sulfide S^{2-}	dihydrogen phosphate H_2PO_4^-
hydrogen carbonate HCO_3^-	sulfite SO_3^{2-}	
hydrogen sulfate HSO_4^-	silicate SiO_3^{2-}	
hydroxide OH^-		
hypochlorite ClO^-		
iodide I^-		
nitrate NO_3^-		
nitrite NO_2^-		
perchlorate ClO_4^-		
permanganate MnO_4^-		

DIATOMIC ELEMENTS: H, N, O, F, Cl, Br, I

S, P, H

AMMONIA GAS: NH_3

Methane Gas: CH_4

Molecular - Covalent

COMMON ACIDS

HCl hydrochloric
 HBr hydrobromic
 $\text{HC}_2\text{H}_3\text{O}_2$ acetic
 HNO_3 nitric
 H_2SO_4 sulfuric
 H_3PO_4 phosphoric

BINARY PREFIXES

1 mono-	6 hexa-
2 di-	7 hepta-
3 tri-	8 octa-
4 tetra-	9 nona-
5 penta-	10 deca-